

Punyashlok Ahilyadevi Holkar Solapur University, Solapur



Name of the Faculty: Science & Technology

CHOICE BASED CREDIT SYSTEM

Syllabus: Chemistry

Name of the Course: B. Sc. Part- III (Sem. V & VI)

(Syllabus to be implemented June 2024)

Punyashlok Ahilyadevi Holkar Solapur University, Solapur
New Choice Based Credit System (CBCS)- (w.e.f. 2024-25)
Structure for B. Sc.-III Chemistry

Subject / Core Course	Name and Type of the Paper		No. of papers/ Practical	Hrs/week			Total Marks Per Paper	UA	CA	Credits
	Type	Name		L	T	P				
Class :	B.Sc.- III Semester – V									
Ability Enhancement Course(AECC)	English (Business English)		Paper- III	4.0	-	--	50	40	10	2.0
Discipline Specific Elective(DSE) (Students can opt any one subjects among the three Subjects excluding interdisciplinary offered at B.Sc II.)	DSE-1A-Physical Chemistry		Paper- IX	3	-	--	100	80	20	4.0
	DSE- 2 A-Inorganic Chemistry		Paper -X	3	-	--	100	80	20	4.0
	DSE- 3 A-Organic Chemistry		Paper- XI	3	-	--	100	80	20	4.0
	ANY ONE from DSE-4A(I) & 4A(II)		Paper- XII	3	-	--	100	80	20	4.0
	DSE-4 A(I)-Analytical and Industrial Physical Chemistry OR DSE-4 A(II)-Methodology and materials of industrial importance									
Grand Total				16	-	--	450	360	90	18
Class :	B.Sc.- III Semester –VI									
Ability Enhancement Course(AECC)	English (Business English)		Paper IV	4.0	-	--	50	40	10	2.0
DSE (Students can opt any one subjects among the three Subjects excluding interdisciplinary offered at B.Sc. II.	DSE- 1B-Physical Chemistry		Paper - XIII	3.0	-	--	100	80	20	4.0
	DSE- 2B-Inorganic Chemistry		Paper- XIV	3.0	-	--	100	80	20	4.0
	DSE- 3B-Organic Chemistry		Paper- XV	3.0	-	--	100	80	20	4.0
	ANY ONE from DSE-4B(I) & 4B(II)		Paper- XVI	3.0	-	--	100	80	20	4.0
	DSE 4B(I)- Analytical and Industrial Organic Chemistry OR DSE 4B(II)-Applied Organic Chemistry									
	SEC									
Total (Theory)				16	-	--	450	360	90	18
DSE - Practical (Annual Exam)	DSE- 1 A& B		Practical- IX & XIII	--	-	5	100	80	20	4.0
	DSE -2 A& B		Practical- X & XIV	--	-	5	100	80	20	4.0
	DSE- 3 A& B		Practical- XI & XV			5	100	80	20	4.0
	DSE- 4 A& B		Practical- XII & XVI			5	100	80	20	4.0
Total (Practicals)						20	400	320	80	16
Grand Total				32		20	1300	1040	260	52

***(Add-on-self learning) – MOOC / SWAYAM COURSE / INTERNSHIP / INDUSTRIAL TRAINING / Courses offered*by College**

Add on College course List should be submitted to the University for Information

General Structure

Theory Examination:

- Structure of B.Sc. course under faculty of science has total 06 semesters for 3years.
- B.Sc.-III comprises of total two semesters (Sem-V and Sem-VI).
Each semester will have Five theory papers (one compulsory English and four Chemistry papers) of 80 marks each (University external examination) and 20 marks for each paper (Internal examination)
The duration of each University theory paper examination will be of 2 hr. and 30 min.
Each theory paper has 20 marks for internal examination. There will be 10 marks unit test and 10 marks home assignment
- At the end of academic year i.e. Semester -VI the practical examination will be conducted.
The weightage of practical is of 320 marks for University external practical examination and 80 i.e (20*4) marks for internal practical examination.

There will be Four theory papers in chemistry of 80 marks for each semester. Their titles and marks distribution are as under (Excluding English).

B Sc –III Sem-V

DSE-1A-Physical Chemistry
DSE- 2A-Inorganic Chemistry
DSE- 3A-Organic Chemistry
DSE 4A(I)- Analytical and Industrial Physical Chemistry
OR
DSE-4 A(II)- Methodology and materials of industrial importance

B Sc- III Sem-VI

DSE- 1B-Physical Chemistry
DSE- 2B-Inorganic Chemistry
DSE- 3B-Organic Chemistry
DSE 4B(I)- Analytical and Industrial Organic Chemistry
OR
DSE-4B(II) Applied Organic Chemistry

Practical Course

Practical Examination will be held at the end of the year.

A) Distribution of marks :

- **Continuous Internal Assessment for chemistry:**

- 1) Practical paper has $20 \times 4 = 80$ marks for internal examination.
- 2) Practical paper has 320 marks for external university practical examination.

There will be four practical courses, one from each Physical, Inorganic, Organic and Analytical practical work.

- 3) The mark distribution of 320 marks for external university practical examination is as follows.

1. Physical Chemistry experiment : 80 marks
2. Inorganic Chemistry experiment : 80 marks
3. Organic Chemistry experiment : 80 marks
4. Analytical Chemistry experiment : 80 marks

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Total marks: 320 marks

Duration of practical examination is three days, six and half hours per day

All answer sheets should be collected at the end of examination.

Practical Marks Distribution

Physical Chemistry experiment	: 80 marks
a) Instrumental	:30
b) Non-instrumental	:30
c) Journal	:10
d) Oral	:10

Inorganic Chemistry experiment	: 80 marks
a) Gravimetric analysis	: 30
b) Volumetric analysis	:20
c) Preparation	:10
d) Journal	:10
e) Oral	:10

Organic Chemistry experiment	: 80 marks
a) Organic Mixture Separation and analysis:	35
b) Preparation	:15
c) Derivative	:10
d) Journal	:10
e) Oral	:10

Analytical Chemistry experiment	: 80 marks
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Experiments from Any two sections	
a) Analytical Physical chemistry	: 30
b) Analytical Inorganic chemistry	: 30
c) Analytical Organic chemistry	: 30
d) Journal	: 10
e) Oral	: 10

CHEMISTRY: Syllabus for B.Sc.-III as per CBCS pattern

SEMESTER –V

PAPER-IX:DSE-1A

PHYSICAL CHEMISTRY

Total Credits:4

Contact hrs: 60

Unit I:

1. Introduction to Quantum Mechanics

[15]

- 1.1 Introduction
- 1.2 Failures of classical mechanics, origin of quantum mechanics
- 1.3 Black body radiation, Stefan-Boltzmann law,
- 1.4 Planck's quantum theory of black body radiation distribution ,
- 1.5 Photoelectric effect, explanation on the basis of quantum theory,
- 1.6 Compton effect
- 1.7 de-Broglie hypothesis
- 1.8 Heisenberg's uncertainty principle (statement explanation)
- 1.9 Schrodinger wave equation- (Derivation not expected)
- 1.10 Physical significance of wave function ψ and ψ^2

Unit II:

2A. Electromotive force

[15]

(Convention : Reduction potentials to be used)

- 2.1 Introduction
- 2.2 Thermodynamics of electrode potentials, Nernst equation for electrode and cell potentials in terms of activities.
- 2.3 Types of electrodes : Description in terms of construction, representation, half cell reaction and emf equation for,
 - i) Metal - metal ion electrode. ii) Amalgam electrode.
 - iii) Metal - insoluble salt electrode. iv) Gas -electrode.
 - v) Oxidation – Reduction electrode.
- 2.4
 - i) Reversible and Irreversible cells.
 - ii) Chemical cells without transference.
 - iii) Concentration cells
 - a. Electrode concentration cell
 - I) Reversible to cation
 - II) Reversible to anion
 - b. Electrolyte concentration cells without transference
- 2.5 Numerical problems.

Unit III:

2B. EMF

(6)

- 2.6 Equilibrium constant from cell emf, determination of the thermodynamic parameters such as ΔG , ΔH and ΔS .
- 2.7 Applications of emf measurements:
 - i) Determination of pH of solution using Hydrogen electrode.
 - ii) Solubility and solubility product of sparingly soluble salts (based on concentration cell).
- 2.8 Numerical problems.

3. Phase Equilibria.

[9]

- 3.1 Introduction
- 3.2 Gibbs phase rule : Phase rule equation and explanation of terms involved in the equation. Phase diagram, true and metastable equilibria.
- 3.3 One component systems : (i) Water system (ii) Sulphur system with explanation for polymorphism.
- 3.4 Two component systems : (i) Eutectic system : (Ag - Pb system); Desilverisation of lead (ii) Formation of compound with congruent melting point (FeCl_3 - H_2O)

Unit IV:

4. Photochemistry

[15]

- 4.1 Introduction
- 4.2 Difference between thermal and photochemical processes.
- 4.3 Laws of photochemistry : Grotthus - Draper law, Lambert law, Lambert - Beer's law (with derivation), Stark – Einstein law.
- 4.4 Quantum yield, Reasons for high quantum yield (e.g. H_2 - Cl_2) and low quantum yield. (e.g. Decomposition of HI and HBr).
- 4.5 Photosensitized reactions - Dissociation of H_2 , Photosynthesis.
- 4.6 Photodimerisation of anthracene.
- 4.7 Jablonski diagram depicting various processes occurring in the excited state :
Qualitative description of fluorescence and phosphorescence.
- 4.8 Chemiluminescence.
- 4.9 Numerical problems.

Reference Books:

1. Physical Chemistry by G. M. Barrow, International student Edition, Mc Graw Hill.
 2. University General Chemistry by C.N.R. Rao, Macmillan.
 3. Physical Chemistry by, R. A. Alberty, Wiley Eastern Ltd.
 4. The Elements of Physical Chemistry by P. W. Atkins, Oxford.
 5. Principles of Physical Chemistry by S. H. Maron, C. H. Prutton, 4th Edition.
 6. Fundamentals of Photochemistry by K. K. Rohatgi-Mukerjee.
 7. Principles of Physical Chemistry by Puri, Sharma, Pathania, Shobhanlal Naginchand and Company, Jalandar.
 8. Text Book of Physical Chemistry by S. Glasstone, Macmillan India Ltd.
 9. Elements of Physical Chemistry by D. Lewis and S. Glasstone (Macmillan).
 10. Principles of Physical Chemistry by Maron and Lando (Amerind).
 11. An Introduction to Electrochemistry by S. Glasstone.
 12. Physical Chemistry by W. J. Moore.
 13. Essentials of Physical Chemistry, Bahl and Tuli (S.Chand).
 14. Quantum Chemistry: R. K. Prasad
 15. Quantum Chemistry: D. A. MacQuerrey
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PAPER –X: DSE-2A
INORGANIC CHEMISTRY

Total Credits: 04
Contact hrs: 60

Unit I.

1. Metal Ligand Bonding in Transition Metal Complexes [15]

A) Crystal Field Theory (CFT).

- 1.1. Introduction - What is CFT?
- 1.2. Basic concept of FT.
- 1.3. Formation of complexes with Crystal field splitting of 'd' orbitals
 - i. Shapes of d orbitals and their electron density region
 - ii. Formation of octahedral Complex with Crystal field splitting of 'd' orbitals, e.g. High spin and low spin octahedral complexes of $\text{Co(III):[CoF}_6\text{]}^{3-}$, $[\text{Co(NH}_3\text{)}_6]^{3+}$.
 - iii. Formation of tetrahedral Complex with Crystal field splitting of 'd' orbitals, e.g. $[\text{CoCl}_4]^{2-}$
 - iv. Formation of square planar Complex with Crystal field splitting of 'd' orbitals e.g. $[\text{Co(CN)}_4]^{2-}$
- 1.4. Jahn–Teller distortion.
- 1.5. Factors affecting the Crystal – field splitting.
- 1.6. Crystal field stabilization energy (Δ): Calculation for octahedral complexes only.
- 1.7. Applications and limitations of CFT.

B) Molecular Orbital Theory (MOT).

- 1.8 Introduction.
- 1.9 Basic concept
- 1.10 Symmetry classes of atomic orbitals
- 1.11 Formation of octahedral complex a) Assumptions b) M.O. energy level diagram for hypothetical octahedral complex.
- 1.12 Examples: octahedral complexes with sigma bonding only such as -e.g. $[\text{Ti(H}_2\text{O)}_6]^{3+}$, $[\text{FeF}_6]^{3-}$, $[\text{Fe(CN)}_6]^{3-}$, $[\text{CoF}_6]^{3-}$, $[\text{Co(NH}_3\text{)}_6]^{3+}$, $[\text{Ni(NH}_3\text{)}_6]^{2+}$
- 1.13 Applications and limitations of MOT.
- 1.14 Comparison between CFT and MOT.

Unit II:

2. Nuclear Chemistry: [15]

- 2.1 Nuclear reaction and energetics of nuclear reactions.
- 2.2 Classification of nuclear reactions and Types of nuclear reactions:
 - i) Artificial transmutation.
 - ii) Artificial radioactivity.
 - iii) Projectile capture reaction.
 - iv) Projectile capture - particle emission reaction.
 - v) Nuclear fission.
 - vi) Nuclear fusion.
- 2.3 Use of Uranium, Thorium and Plutonium for: a. Nuclear reactor b. Atomic Bomb.
- 2.4 Applications of radioisotopes as tracers.
 - i) Chemical investigation - Esterification.
 - ii) Structural determination – Phosphorus pentachloride.

- iii) Analytical Chemistry - isotopic dilution method for determination of volume of blood.
- iv) Age determination - Dating by ^{14}C .

Unit III:

3. Bioinorganic Chemistry:

[15]

- 3.1 Essential and trace elements in biological process.
 - i) Essential elements a) Macro / major elements b) Micro/trace/minor elements
 - ii) Non-essential elements
- 3.2 Metalloporphyrins with special reference to haemoglobin and myoglobin.
 - i) Structure of Haemoglobin (Hb)
 - ii) Structure of Myoglobin (Mb)
 - iii) Function of Haemoglobin (Hb) and Myoglobin (Mb) as Oxygen transport from lungs to tissues
 - iv) Function of Haemoglobin as Carry back CO_2 to lungs
 - v) Co-operativity
 - vi) Oxygen binding curve
 - vii) Difference between Haemoglobin (Hb) and Myoglobin (Mb)
- 3.3 Role of alkali and alkaline earth metal ions with special reference to Na^+ , K^+ and Ca^{2+} .
 - i) Role of Na^+ and K^+
 - ii) Role of Ca^{2+} .

Unit IV:

4. Catalysis

[07]

- 4.1 Introduction
- 4.2 Classification of catalytic reactions : Homogeneous & Heterogeneous
- 4.3 Types of catalysis
- 4.4 Characteristics of catalytic reactions
- 4.5 Mechanism of catalysis:
 - i) Intermediate compound theory
 - ii) Adsorption theory.
- 4.6 Industrial Applications of Catalysis.

5. Fertilizers

[08]

- 5.1 Nutrient Functions in plant growth:
 - Nitrogen, Phosphorous, Potassium, Calcium, Magnesium, Sulphur, Boron, Iron, Zinc, Manganese, Copper, Molybdenum, Chlorine, Role of these nutrients as : Functions, Excess supply and Deficiency.
- 5.2 Definition and qualities of an ideal fertilizers:
- 5.3 Classification or types of fertilizers:
- 5.4 Manufacture of fertilizers, eg. Urea, Ammonium sulphate, Superphosphate, Triple superphosphate, Ammonium phosphate.
- 5.5 Mixed fertilizers, Compound or complex fertilizers.
- 5.6 Pollution caused by fertilizers

Reference Books:

1. Concise Inorganic Chemistry (ELBS, 5th Edition) - J. D. Lee.
2. Inorganic Chemistry (ELBS, 3rd Edition) D. F. Shriver, P. W. Atkins, C. H. Langford, Oxford University Press, 2nd Edition.
3. Inorganic Chemistry (Harper International, 3rd edition) J. E. Huheey Harper and Row.
4. Basic Inorganic Chemistry : Cotton and Wilkinson.
5. Advanced Inorganic Chemistry (4th Edn.) Cotton and Wilkinson.
6. Concepts and Models of Inorganic Chemistry : Douglas and Mc. Daniel. 3rd Edition. John Wiley publication.
7. Fundamental concepts of Inorganic Chemistry by E. S. Gilreath.
8. Structural principles in inorganic compounds. W. E. Addison.
9. T. B. of Inorganic analysis - A. I. Vogel.
10. Theoretical principles of Inorganic Chemistry - G. S. Manku.
11. Theoretical Inorganic Chemistry by Day and Selbina.
12. Co-ordination compounds- S F A Kettle.
13. Modern Aspects of Inorganic Chemistry. E. Sharpe.
14. New guide to Modern Valence Theory by G. I. Brown.
15. Essentials of Nuclear Chemistry by H. J. Arnikar.
16. Organometallic Chemistry by R. C. Mahrotra A. Sing, Wiley Eastern Ltd. New Delhi.
17. Inorganic Chemistry by A. G. Sharpe, Addison - Wesley Longman -Inc.
18. Principles of Inorganic Chemistry by Puri, Sharma and Kalia, Vallabh Publication. Pitampur Delhi.
19. Text book of Inorganic Chemistry by K. N. Upadhyaya Vikas Publishing House –New Delhi.
20. Progress in inorganic polymer by Laport and Leigh.
21. Co-ordination compounds by Baselo and Pearson.
22. Advanced inorganic chemistry, Vol. I and II Satyaprakash, G. D. Tuli, S. K. Basu and Madan (S. Chand)
23. Selected Topics in inorganic chemistry by W U Malic, G. D. Tuli, R. D. Madan. (S. Chand)
24. Industrial chemistry part I and II by A. K. De
25. Industrial Chemistry, By – B K Sharma, Goel Publishing House 16th Edition

Paper – XI: DSE-3A
Organic Chemistry

Total Credits: 4

Contact hrs: 60

Unit I:

1. Spectroscopic Methods.

15

A) Infrared Spectroscopy

[07]

- 1.1 Introduction.
- 1.2 Principle of IR spectroscopy.
- 1.3 Double beam IR spectrophotometer- Schematic diagram.
- 1.4 Fundamental modes of vibrations.
- 1.5 Types of vibrations.
- 1.6 Hooke's law.
- 1.7 Factors affecting values of vibrational frequencies.
- 1.8 Conditions for absorption of radiation and selection rule.
- 1.9 Fundamental group regions of IR spectrum.
- 1.10 Functional group region, Finger print region, Aromatic region.
- 1.11 Characteristic absorption of various functional groups.
- 1.12 Applications of IR spectroscopy – Determination of structure, Identification of functional groups, spectral problems based on IR.

B) Mass spectroscopy

[08]

- 1.13 Introduction.
- 1.14 Theory of mass spectroscopy
- 1.14 Mass spectrometer – schematic diagram
- 1.15 Formation of ions by ionization
- 1.16 Types of ions with examples.
- 1.17 Applications of mass spectroscopy.
 - i) Determination of molecular weight.
 - ii) Determination of molecular formula.

Unit II:

2. NMR Spectroscopy

[15]

- 2.1 Introduction.
- 2.2. Proton magnetic resonance (^1H) spectroscopy (PMR).
- 2.3 Principles of PMR spectroscopy.
- 2.4 Magnetic and non-magnetic nuclei.
- 2.5. Theory of PMR spectroscopy - spinning nuclei, magnetic moment and magnetic field, precessional motion of nuclei without mathematical details, nuclear resonance.
- 2.6 NMR - Instrument. Schematic diagram.
- 2.7 Shielding and deshielding effect.
- 2.8 Chemical shift, measurement of chemical shift by delta scale and tau scale.
- 2.9 TMS as reference. Advantages of TMS.
- 2.10 Peak area (integration).
- 2.11 Spin - spin splitting ($n + 1$ rule).
- 2.12 Definition of coupling constant (J value) of first order coupling.
- 2.13 PMR spectra of ethanol, ethyl bromide, acetaldehyde, 1, 1, 2 - tribromoethane, ethyl acetate, acetophenone, benzaldehyde, propanoic acid and benzoic acid.
- 2.14 Problems pertaining to the structure elucidation of simple organic compounds using PMR spectroscopic data (supporting IR and UV data to be given).

Unit III.

[15]

3. Stereochemistry

[08]

- 3.1 Introduction.
- 3.2 Baeyer's strain theory.
- 3.3 Theory of strainless rings.
- 3.4 Conformation and stability of cyclohexane and monosubstituted cyclohexanes– methyl cyclohexane.
- 3.5 Locking of conformation in t-butyl cyclohexane.
- 3.6 Stereoselective and stereospecific reactions:
 - i) Stereochemistry of addition of halogens to alkenes: syn- and anti-addition.
Example -Addition of bromine to 2-butene. (mechanism not expected)
 - ii) Alkaline hydrolysis of 2-chlorobutane to 2-butanol (Example of S_N^2 reaction)

4. Name reactions

[07]

Mechanism and applications of following reactions:

- 4.1 Stobbe condensation.
- 4.2 Oppenauer oxidation.
- 4.3 Meerwein Ponndorf Verley reduction.
- 4.4 Reformatsky reaction.
- 4.5 Wagner –Meerwein Rearrangement.
- 4.6 Hofmann rearrangement reaction.
- 4.7 Wittig reaction.
- 4.8 Related problems.

Unit IV:

5. Organic synthesis via Enolates

[15]

- 5.1 Introduction - Reactive methylene group.
- 5.2 Ethyl acetoacetate - synthesis by Claisen condensation, acidity of methylene hydrogen (salt formation), Keto-enol tautomerism, synthetic applications – Synthesis of alkyl and dialkyl derivatives, monobasic, dibasic and α - β -unsaturated acid, heterocyclic compound.
- 5.3 Diethyl malonate - Synthesis, acidity of methylene hydrogen (salt formation).
Synthetic applications - Synthesis of alkyl and dialkyl derivatives, monobasic, dibasic acid, α - β - unsaturated acid, α -amino acid and heterocyclic compound.

Reference Books:

- 1) Organic Chemistry: D. J. Cram and G. S. Hammond, McGraw Hill book Company, New York.
- 2) Organic Chemistry: I. L. Finar, The English Language Book Society, London.
- 3) A Guide Book to mechanism in Organic Chemistry: Peter Sykes, Longman Greenand Co. Ltd. London 6thEdition.
- 4) Organic Chemistry: R. T. Morrison and R. N. Boyd, Prentice Hall of IndiaPrivate Limited, New Delhi. 6thEdition.
- 5) Text book of organic Chemistry: L. N. Ferguson, N. D. Van Nostrand Company Indian Edition, Affiliated East west press private Ltd. New Delhi.
- 6) Organic Chemistry Vol. I, II and III: S. M. Mukharji, S. P. Singh, R. P. Kapoor Wiley Eastern, Limited, New Delhi.
- 7) A text book of organic Chemistry: K. S. Tewari, S. N. Mehrotra, N.K. Vishnoi Vikas Publishing House Private Ltd. New Delhi.
- 8) A text book of Organic Chemistry: Arun Bahl and B. S. Bahl , S. Chandand Company Ltd. 6thEdition.
- 9) Heterocyclic Chemistry Synthesis, Reactions and Mechanism: Raj K. Bansal, Wiley Easter Ltd., New Delhi.
- 10) Reaction Mechanism and Reagents in Organic Chemistry: G. R. Chatwal, Himalaya Publishing House, New Delhi.
- 11) Stereochemistry conformation and mechanism: P. S. Kalsi, New Age International Publishers, 4th Edition.
- 12) Organic Chemistry Volume I and II: I. L. Finar ELBS with Longman 6th Edition.
- 13) Organic Chemistry Volume I and II : William Kemp, ELBS with Mc. Million 3rd Edition.
- 14) Advanced Organic Chemistry: Jerry March, Wiley Eastern Ltd.
- 15) Spectroscopy of Organic compounds: P. S. Kalsi.
- 16) Modern Methods of Organic Synthesis, W Carruthers, Iain Coldham, Cambridge University Press
- 17) Organic Chemistry: Fieser and Fieser.
- 18) Principles of Organic Chemistry: English and Cassidy.
- 19) Elementary Organic Absorption Spectroscopy: Y. R. Sharma.
- 20) Spectroscopy: V. M. Parikh.
- 21) Stereochemistry of Carbon Chemistry: Eliel.
- 22) Principles of Organic Chemistry: M. K. Jain.
- 23) Organic Chemistry by Clayden, Greeves, Warren and Wothers Oxford press.
- 24) Organic Chemistry: A Comprehensive degree text and source book by Hanes Baeyers and Wolfgang Walter Albion Chemical Science Series.
- 25) Reactions, Rearrangements and reagents: S.N. Sanyal, Bharati Bhawan publishers and Distributors Patna.

PAPER-XII : DSE-4A(I)
ANALYTICAL AND INDUSTRIAL PHYSICAL CHEMISTRY

Total Credits: 4

Contact hrs: 60

Unit. I: [15]

1. Colorimetry [7]

1.1 Introduction

1.2 General discussion of theory of colorimetry : Lambert law, Beer's law (Derivation not expected), Terms used in Colorimetry, Application of Beer's law, Deviation from Beer's law.

1.3 Classification of methods of color measurement or comparison, Photoelectric photometer method - single cell photo-electric colorimeter.

2. Flame photometry [8]

2.1 General principles.

2.2 Instrumentation : Block diagram,

Burners: Total consumption burner, premix or laminar-flow burner and Lundergraph burner,

Mirrors, Slits, Monochromators, Filters and Detectors.

2.3 Applications in qualitative and quantitative analysis.

2.4 Limitations of flame photometry.

Unit II.

3. Potentiometry [15]

3.1. Introduction.

3.2. Detail study of calomel, quinhydrone and glass electrodes and their use in determination of pH.

3.3. Basic circuit diagram of direct reading potentiometer

3.4 Potentiometric titrations : Classical and analytical methods for locating endpoints,

i) Acid – Base titrations.

ii) Redox - titrations.

iii) Precipitation titrations.

3.5 Advantages of potentiometric titrations.

Unit III.

4. Electroplating [15]

4.1 Introduction.

4.2 Electrolysis, Faraday's laws, Cathode current efficiency.

4.3 Basic principles of electroplating, cleaning of articles.

4.4 Electroplating of Nickel and Chromium.

4.5 Anodising.

Unit IV.

5. Conductometry: [15]

5.1 Basic circuit of D.C. Wheatstone bridge, Measurement of conductance by Wheatstone bridge, use of alternating current, conductivity water, Different types of conductivity cells, cell constant and its determination.

5.2 Experimental determination of specific, equivalent and molar conductance.

5.3 Conductometric acid-base titrations

i. Strong acid against strong base

ii. Strong acid against weak base

iii. Weak acid against strong base.

iv. Weak acid against weak base.

5.4 Advantages of conductometric titrations

Reference Books :

1. Text book of Quantitative Inorganic Analysis - By A. I. Vogel (ELBS and Longman 3rd Edition).
 2. Instrumental methods of Chemical analysis by Willard, Merit and Dean.
 3. Instrumental methods of Chemical analysis by Chatwal and Anand (Himalaya Publication).
 4. Principles of electroplating and eletroforming by Blum and Hogaboom, Mac Graw - Hill Book Co. 3rd Edn.
 5. Vogel's text book of Quantitative Inorganic Analysis by Bassett and Denny etc. ELBS and Longman 4th Edition.
 6. Principles of Physical Chemistry by Puri, Sharma, Pathania, Shobhanlal Nagin Chand and Company, Jalandar.
 7. Text Book of Physical Chemistry by S. Glasstone, McMillan India Ltd.
 8. Elements of Physical Chemistry by D. Lewis and S. Glasstone (McMillan).
 9. Principles of Physical Chemistry by Maron and Lando (Amerind).
 10. An Introduction to Electrochemistry by S. Glasstone.
 11. Physical Chemistry by W. J. Moore.
 12. Essentials of Physical Chemistry, Bahl and Tuli (S. Chand).
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PAPER-XII : DSE-4A(II)
METHODOLOGY AND MATERIALS OF INDUSTRIAL IMPORTANCE

Total Credits: 4

Contact hrs: 60

Unit I.

1. Data Analysis (15)

- 1.1 The Investigative Approach: Making and Recording Measurements. SI Units and their use. Scientific method and design of experiments.
- 1.2 Analysis and Presentation of Data: Descriptive statistics. Choosing and using statistical tests.
- 1.3 Chemometrics. Analysis of variance (ANOVA), Correlation and regression, Curve fitting, fitting of linear equations, simple linear cases, weighted linear case, analysis of residuals,
- 1.4 General polynomial fitting, linearizing transformations, exponential function fit, 'r' and its abuse.
- 1.5 Basic aspects of multiple linear regression analysis.

Unit II.

2. Chemical Safety and Ethical Handling of Chemicals (15)

- 2.1 Safe working procedure and protective environment, protective apparel, emergency procedure and first aid, laboratory ventilation.
- 2.2 Safe storage and use of hazardous chemicals,
- 2.3 Procedure for working with substances that pose hazards, flammable or explosive hazards,
- 2.4 Procedures for working with gases at pressures above or below atmospheric – safe storage and disposal of waste chemicals, recovery, recycling and reuse of laboratory chemicals,
- 2.5 Procedure for laboratory disposal of explosives, identification, verification and segregation of laboratory waste, disposal of chemicals in the sanitary sewer system.
- 2.6 Incineration and transportation of hazardous chemicals.

Unit III.

3. Nanomaterials (15)

- 3.1 Overview of nanostructures and nanomaterials: classification.
- 3.2 Preparation of gold and silver metallic nanoparticles, self-assembled nanostructures-control of nanoarchitecture-one dimensional control.
- 3.3 Carbon nanotubes and inorganic nanowires.
- 3.4 Bio-inorganic nanomaterials, DNA and nanomaterials, natural and antisocial nanomaterials,
- 3.5 Bio-nanocomposites.

Unit IV.

4. Composite materials (15)

- 4.1 Introduction, limitations of conventional engineering materials, role of matrix in composites,
- 4.2 Classification, matrix materials, reinforcements, metal-matrix composites, polymer-matrix composites, fibre-reinforced composites.
- 4.3 Environmental effects on composites.
- 4.4 Applications of composites.

Reference Books

- 1. Practical skills in chemistry, Dean, J. R., Jones, A. M., Holmes, D., Reed, R., Weyers, J. & Jones, A. (2011) 2nd Ed. Prentice-Hall, Harlow.
- 2. Data analysis for chemistry, Hibbert, D. B. & Gooding, J. J. (2006) Oxford University Press.
- 3. Errors of observation and their treatment, Topping, J. (1984). Fourth Ed., Chapman Hall, London.
- 4. Quantitative chemical analysis, Harris, D. C. 6th Ed., Freeman (2007) Chapters 3-5.
- 5. How to use Excel in analytical chemistry and in general scientific data Analysis, Levie, R. de, Cambridge Univ. Press (2001) 487 pages.
- 6. Chemical safety matters – IUPAC – IPCS, Cambridge University Press, 1992.
- 7. Inorganic Solids: An introduction to concepts in solid-state structural Chemistry, Adam, D.M. John Wiley & Sons, 1974.
- 8. Introduction to Nanotechnology, Poole, C.P. & Owens, F.J. John Wiley & Sons, 2003.

SEMESTER-VI
PAPER-XIII:DSE-IB
PHYSICAL CHEMISTRY

Total Credits: 4
Contact hrs:60

Unit I.

1. Spectroscopy

[15]

- 1.1 Introduction
- 1.2 Electromagnetic radiation.
- 1.3 Electromagnetic spectrum, Energy level diagram.
- 1.4 Rotational spectra of diatomic molecules : Rigid rotor model; moment of inertia (derivation not expected); energy levels of rigid rotor, selection rule; spectral intensity; distribution using population distribution (Maxwell - Boltzmann distribution), determination of bond length; isotope effect. Interaction of radiation with rotating molecule.
- 1.5 Vibrational spectra of diatomic molecules: Simple Harmonic oscillator model, Vibrational energies of diatomic molecules, Determination of force constant, zero point energy. The Anharmonic oscillator, overtones and hot band. Interaction of radiation with vibrating molecules.
- 1.6 Raman spectroscopy: Introduction, Rayleigh scattering. Raman Scattering, classical theory of Raman effect and quantum theories of Raman effect. Polarization of light and the Raman effect. Mutual exclusion principle.
- 1.7 Numerical problems.

Unit II.

2. Solutions

[15]

- 2.1 Introduction
- 2.2 Ideal solutions, Raoult's law, vapour pressure of ideal and non ideal solutions of miscible liquids.
- 2.3 Vapour pressure and boiling point diagrams of miscible liquids.
 - Type I : Systems with intermediate total vapour pressure. (i.e. System in which B.P. increases regularly - Zeotropic)
 - Type II : Systems with a maximum in the total vapour pressure. (i.e. System with a B.P. minimum - Azeotropic)
 - Type III : Systems with a minimum in the total vapour pressure. (i.e. System with a B.P. Maximum - Azeotropic)
- 2.4 Distillation of miscible liquid pairs.
- 2.5 Solubility of partially miscible liquids.
 - (i) Maximum solution temperature type : Phenol – water system.
 - (ii) Minimum solution temperature type : Triethyl amine – water system.
 - (iii) Maximum and minimum solution temperature type : Nicotine – water system.

Unit III.

3. Thermodynamics

[15]

- 3.1 Introduction
- 3.2 Free energy : Gibbs function (G) and Helmholtz function (A), Criteria for thermodynamic equilibrium and spontaneity.
- 3.3 Relation between G and H : Gibbs Helmholtz equation.
- 3.4 Phase equilibria : Clapeyron – Clausius equation.
- 3.5 Thermodynamic derivation of law of mass action, van't Hoff isotherm and isochore.
- 3.6 Fugacity and activity concepts.
- 3.7 Numerical problems.

Unit IV.

4. Chemical Kinetics

[15]

- 4.1 Introduction, simultaneous reactions such as opposing reactions, side reactions, consecutive reactions and chain reactions. [Derivations of rate Equations for these reactions are not expected.]
- 4.2 Effect of temperature on the rate of reaction.
 - 1. Temperature coefficient
 - 2. Arrhenius equation
 - 3. Energy of activation
- 4.3 Theories of reaction rate:

1. Collision theory and
 2. Transition state theory
- 4.4 Third order reaction with equal concentration of all reactants, their characteristics and examples
- 4.5 Numerical problems.

Reference Books :

1. Principles of Physical Chemistry by Maron and Pruton 4th edition.
2. Principles of Physical Chemistry by Puri, Sharma, Pathania, Shobhanlal Nagin chand and Company, Jalandar.
3. Text Book of Physical Chemistry by S. Glasstone, McMillan India Ltd.
4. Elements of Physical Chemistry by D. Lewis and S. Glasstone (McMillan).
5. Principles of Physical Chemistry by Maron and Lando (Amerind).
6. Thermodynamics for chemists by S Glasstone.
7. Physical Chemistry by W. J. Moore.
8. Essentials of Physical Chemistry, Bahl and Tuli (S.Chand).
9. Basic Chemical Thermodynamics by V V Rao (McMillan)
10. An introduction to chemical thermodynamics by R. R. Mishra and R. P. Rastogi.
11. Fundamentals of molecular spectroscopy by C. N. Banwell and McCash- Tata McGraw Hill

PAPER-XIV: DSE-2B
INORGANIC CHEMISTRY

Total Credits: 04
Contact hrs: 60

Unit I:

1. Study of F-block Elements

[15]

1.1 Lanthanides:-

- I) Introduction
- II) Electronic configuration
- III) Occurrence
- IV) Separation of Lanthanides
 - i) Bulk separation methods
 - ii) Individual separation of lanthanides- Mention names of methods only (Ion exchange method in detail)

1.2 Actinides:-

- I) Introduction
- II) Electronic configuration
- III) General Methods of preparation–
 - a. Neutron-capture followed by β -decay
 - b. Accelerated projectile bombardment method
 - c. Heavy-ion bombardment method

Unit II:

2. Metals and Semiconductors

[15]

2.1 Introduction.

2.2 Properties of metallic solids.

2.3 Theories of bonding in metal.

- a) Free electron theory.
- b) Molecular orbital theory (Band theory).

2.4 Classification of solids as conductor, insulators and semiconductors on the basis of band theory.

2.5 Semiconductors:

- a) Types of semiconductors - intrinsic and extrinsic semiconductors.
- b) Applications of semiconductors.

2.6 Superconductors:

- a) Ceramic superconductors - Preparation and structures of mixed oxide $\text{YBa}_2\text{Cu}_3\text{O}_{7-x}$
- b) Applications of superconductors.

Unit III:

3. Structural Chemistry

[15]

3.1 Structural study of following compounds.

- i) Diborane.
- ii) Borazine.
- iii) Xenon compounds $\rightarrow \text{XeF}_2, \text{XeF}_6, \text{XeO}_4$ (w.r.t. VBT only.)

3.2 Structural study of Oxides of Sulphur and Phosphorous:

- i) Oxides of Sulphur : SO_2 and SO_3
- ii) Oxides of Phosphorous : P_4O_6 and P_4O_{10}

Unit IV:

[15]

4. Corrosion and Passivity.

[07]

Corrosion:-

4.1 Introduction

4.2 Types of corrosion

4.3 Electrochemical theory of corrosion

4.4 Factors affecting the corrosion

- i) Position of metal in emf series.
- ii) Purity of metal.
- iii) Effect of moisture.

- iv) Effect of oxygen.
- v) Hydrogen overvoltage.

4.5 Methods of protection of metals from corrosion.

Passivity:-

- 4.6 Definition.
- 4.7 Types of passivity.
- 4.8 Oxide film theory.
- 4.9 Application of passivity.

5. Organometallic Chemistry.

[08]

- 5.1 Introduction -Definition,
- 5.2 Nomenclature of organometallic compounds.
- 5.3 Synthesis and structural study of alkyl and aryl compounds of Li, Be and Al.
- 5.4 Mononuclear carbonyl and nature of bonding in simple metal carbonyls.

Reference Books :

1. Concise Inorganic Chemistry (ELBS, 5th Edition) - J. D. Lee.
2. Inorganic Chemistry (ELBS, 3rd Edition) D. F. Shriver, P. W. Atkins, C. H. Lang Ford, Oxford University Press, 2nd Edition.
3. Inorganic Chemistry (Harper International, 3rd edition) J. E. Huheey Harper and Row.
4. Basic Inorganic Chemistry : Cotton and Wilkinson.
5. Advanced Inorganic Chemistry (4th Edn.) Cotton and Wilkinson.
6. Concepts and Models of Inorganic Chemistry : Douglas and Mc. Daniel. 3rd Edition. John Wiley publication.
7. Fundamental concepts of Inorganic Chemistry by E. S. Gilreath.
8. Structural principles in inorganic compounds. W. E. Addison.
9. T. B. of Inorganic analysis - A. I. Vogel.
10. Theoretical principles of Inorganic Chemistry - G. S. Manku.
11. Theoretical Inorganic Chemistry by Day and Selbin.
12. Co-ordination compounds S F A Kettle.
13. Modern Aspects of Inorganic Chemistry. E. Sharpe.
14. New guide to Modern Valence Theory by G. I. Brown.
15. Essentials of Nuclear Chemistry by H. J. Arnikar.
16. Organometallic Chemistry by R. C. Mahrotra A. Sing, Wiley Eastern Ltd. New Delhi.
17. Inorganic Chemistry by A. G. Sharpe, Addison - Wesley Longman -Inc.
18. Principles of Inorganic Chemistry by Puri, Sharma and Kalia, Vallabh Publication. Pitampur Delhi.
19. Text book of Inorganic Chemistry by K. N. Upadhyaya Vikas Publishing House – New Delhi.
20. Progress in inorganic polymer by Laport and Leigh.
21. Co-ordination compounds by Baselo and Pearson.
22. Organometallic Chemistry by P. L. Pauson.
23. Advanced inorganic chemistry, Vol. I and II Satyaprakash, G. D. Tuli, S. K. Basu and Madan
24. Selected Topics in inorganic chemistry by W U Malik, G. D. Tuli, R. D. Madan. (S. Chand)
25. Industrial chemistry part I and II by A. K. De
26. Industrial chemistry by B. K. Sharma

Unit I.

1. Heterocyclic compounds

[15]

1.1 Introduction and classification

1.2 Pyrrole

1.2.1 Methods of synthesis

i) From acetylene

ii) From furan

iii) From succinamide

1.2.2 Physical properties

1.2.3 Reactivity of pyrrole

i) Basic character

ii) Acidic character

iii) Electrophilic substitution with general mechanism

1.2.4 Chemical reactions

i) Reduction

ii) Oxidation

iii) Nitration

iv) Sulphonation

v) Halogenation

vi) Friedel Craft's reaction

vii) Coupling reaction

1.3 Pyridine

1.3.1 Methods of synthesis

i) From acetylene and hydrogen cyanide

ii) From piperidine

1.3.2 Physical properties

1.3.3 Chemical reactions

i) Basic character

ii) Electrophilic substitution reactions : Nitration, Sulphonation and Bromination

iii) Nucleophilic substitution - General mechanism, Reactions with sodamide, sodium hydroxide and n-Butyllithium.

1.4 Quinoline

1.4.1 Synthesis - Skraup's synthesis

1.4.2 Physical properties.

1.4.3 Reactions of quinoline

i) Electrophilic substitution reactions - Nitration and sulphonation.

ii) Nucleophilic substitution reactions – Reactions with sodamide, alkyl lithium and aryl lithium

iii) Reduction

Unit II:

2. Carbohydrates

[15]

2.1 Introduction

2.2 Classification and nomenclature

2.3 Monosaccharide D-glucose - Open chain structure

2.4 Chain lengthening of Aldoses –Kiliani synthesis

2.5 Chain shortening of Aldoses - Weerman's reaction

2.6 Interconversion of glucose and fructose

2.7 Configuration of D-glucose from D-arabinose

2.8 Objections against open chain structure of D-glucose.

2.9 Mutarotation with mechanism.

2.10 Ring structure of D-glucose - Determination of size of ring by

i) Methylation method.

2.1 Disaccharides - Introduction, sucrose and lactose - sources, structural formulae and uses.

2.12 Polysaccharides–Introduction, Starch and Cellulose - sources, structural formulae and uses

Unit III. [15]

3. Vitamins and Hormones [06]

3.1 General idea of vitamins, structure and synthesis of vitamin A

3.2 General idea of hormones, structure and synthesis of Adrenaline and Thyroxin

4. Pharmaceuticals [09]

4.1 Introduction

4.2 Qualities of ideal drug

4.3 Methods of classification of drugs - Classification based on the therapeutical action

4.4 Brief idea of penicillin-G (constitution, synthesis not expected)

4.5 Synthesis and uses of the following drugs:

- i) Antimalarials -Paludrin
- ii) Antituberculars - Isoniazide and Ethambutol
- iii) C. N. S. drugs -Phenobarbitone
- iv) Antidiabetics -Tolbutamide
- v) Anti-inflammatory drugs -Ibuprofen
- vi) Antibiotics -Chloromycetin
- vii)Anticancer drugs : Chlorambucil (Leukeran)

Unit IV. [15]

5. Synthetic dyes [08]

5.1 Introduction, Qualities of good dye

5.2. Classification based on constitution and methods of applications

5.3 Witt's theory - Colour and constitution

5.4 Synthesis of Orange IV, Malechite green, phenolphthalein

6. Agrochemicals [07]

6.1 General idea of agrochemicals including pyrethroids.

6.2 Synthesis and uses of the following agrochemicals:

- i) Indole-3-acetic acid.
- ii) Monocrotophos
- iii) Methoxychlor
- iv) Ethoppan
- v) Carbaryl
- vi) Baygon

Reference Books :

- 1) Organic Chemistry - Cram D. J. and Hammond G.S. McGraw Hill book Company New York.
- 2) Organic Chemistry - Finar I. L. The English Language Book Society, London.
- 3) A Guide Book to mechanism in Organic Chemistry - Peter Sykes Longman Greenand Co. Ltd. London 6thEdition.
- 4) Organic Chemistry - R. T. Morrison and R. N. Boyd Prentice Hall of India private limited New Delhi. 6thEdition.
- 5) Text book of organic Chemistry - Ferguson L. N. D. Van Nostrand Company Indian Edition, Affiliated East West press private Ltd. New Delhi.
- 6) Organic Chemistry Vol. I, II and III - S. M. Mukharji, S. P. Singh, R. P. Kapoor Wiley Estern, Limited, New Delhi.
- 7) A text book of organic Chemistry - K. S. Tewari, S. N. Mehrotra, N. K. Vishnoi Vikas Publishing House Private Ltd. NewDelhi.
- 8) A text book of Organic Chemistry - ArunBahl and B. S. Bahl S. Chand and Company Ltd. 6thEdition.
- 9) Heterocyclic Chemistry Synthesis, Reactions and Mechanism - Raj K. Bansal Wiley Easter Ltd. New Delhi.
- 10) Reaction Mechanism and reagents in Organic Chemistry - G. R. Chatwal Himalaya Publishing House New Delhi.
- 11) Organic Chemistry Volume I and II - I. L. Finar ELBS with Longman 6thEdition.
- 12) Organic Chemistry Volume I and II - William Kemp ELBS with Macmillan 3rd Ed.
- 13) Advanced Organic Chemistry - Jerry March Wiley Eastern Ltd.
- 14) Organic Chemistry – Fieser and Fieser.
- 15) Principles of Organic Chemistry - English and Cassidy.
- 16) Chemicals for crop improvement and pest management - Green, Hartly and West.
- 17) Chemistry of pesticides - K. H. Buchel (T.W.).
- 18) Medical Chemistry -Burger.

- 19) Principles of Organic Chemistry - M. K. Jain.
- 20) Organic Chemistry by Clayden, Greeves, Warren and Wothers Oxford press.
- 21) Organic Chemistry - A Comprehensive degree text and source book by Hanes Baeyers and Wolfgang Walter Albion Chemical Science Series.
- 22) Reactions, Rearrangements and reagents - S.N. Sanyal, Bharati Bhawan publishers and Distributors Patna.
- 23) Synthetic Organic Chemistry-Kamlesh Bansal.
- 24) Synthetic Organic Chemistry-GurudeepChatwal.
- 25) Chemistry of Insecticides – U.S. SreeRamulu.
- 26) Medicinal Chemistry- Ashitosh Kar.

Paper-XVI : DSE-4B(I)
Analytical and Industrial Organic Chemistry

Total Credits: 4
Contact hrs: 60

Unit I:

1. Soaps and Detergents

[15]

Soaps

- i) Raw materials
- ii) Types of soaps
- iii) Manufacture of soap – Hot process
- iv) Cleansing action of soaps

Detergents

- i) Raw materials
 - ii) Types of detergents - Cationic, anionic, amphoteric, neutral detergents
 - iii) Preparation of teepol and deriphat
- Comparison between soaps and detergents

Unit II.

2. Synthetic Polymers

[15]

Introduction

Classification:

- i) According to origin, composition, method of preparation and general physical properties
- ii) Classification based upon structure

Process of addition polymerisation - free radical polymerisation of alkenes and Dienes

Ionic polymerisation

Ziegler – Natta polymerisation

Methods of preparation and uses of:

- i) Polystyrene ii) PVC iii) Phenol formaldehyde resin iv) Polyurethane

Natural rubber : General idea and vulcanisation

Synthetic rubbers : Synthesis and uses of:

- i) Polychloroprene ii) Buna rubber - Buna N and Buna S

Unit III.

3. Sugar and Alcohol Industry

[09]

Manufacture of raw cane sugar

Refining of raw sugar

White sugar

By-products of sugar industry

Manufacture of ethyl alcohol from molasses

Rectified spirit, Denatured spirit absolute alcohol and power alcohol

By-products of alcohol industry

4. Green Chemistry

[06]

Introduction - Twelve principles of green chemistry

PTC: Introduction, Role in organic reactions catalysis

Biocatalytic reactions - Hydroxylation and oxidation using enzymes

Introduction to microwave assisted reactions

Ionic liquids – Introduction and examples of ionic liquids

Unit IV:

5. Synthetic Reagents

[08]

Sodium borohydride: Use in reduction of aldehydes and ketones

Lithium aluminium hydride: Use in reduction of aldehydes, ketones, acids, amides and esters

Osmium tetroxide : Hydroxylation of alkenes

1,3-dithiane : Umpolung concept, reactions with alkyl halide and acyl halide

Selenium dioxide : Oxidation of carbonyl compounds and allylic oxidation

6. Chromatography

[07]

Introduction

General principles

Study of following chromatographic techniques with reference to principle,

methodology and applications

- i) Column chromatography
- ii) Thin layer chromatography
- iii) Gas chromatography

Reference) Books:

1. Basic Concepts of Analytical Chemistry - S. M. Khopkar, Wiley Eastern Ltd. Bombay.
2. Industrial Chemistry - R. K. Das, Asia Publishing, Mumbai.
3. Text Book of Quantitative Organic Analysis - A. I. Vogel, Pearson Edn. Delhi.
4. Quantitative Organic Chemistry - A. I. Vogel, Pearson Edn. Delhi.
5. Hand Book of Organic Analysis - H. T. Clarke, Arnold Heinemann Pub. Delhi.
6. Advanced Organic Chemistry - B. S. Bahl and Arun Bahl, S. Chand Comp. Delhi.
7. Riegel's Handbook of Industrial Chemistry - J. A. Kent, Van. Nostrand, London.
8. Chemical Process Industries - Shreve and Brinic - Ostin, Magraw Hill, NewYork.
9. Analytical Chemistry- Walton.
10. Biotechnology and Applied Microbiology - Alani andMoo-Young.
11. Immobilize Biocatalysis– JoyWleser.
12. Introduction to Polymer Chemistry - Raymond B.Seymour.
13. Polymer Science - V. R. Gowarikar, N. V. Viswanathan and JayadevSreedhar
Willey EasternLimited.
14. Advances in Green Chemistry: Chemical synthesis using MW-irradiation by R. S.Varma.
15. GreenChemistry:EnvironmentFriendlyalternatives-RashmiSanghiandM.
M. Srivastava (Eds) (c) 2003 Narosa Publishing House, New Delhi, India.
16. Reactions, rearrangements and reagents : S. N.Sanyal
17. Organic reaction mechanism : V. K. Ahluwalia and K.R.KParashar
18. Environment friendly synthesis using ionic liquids: JairtonDupont,
Toshiyuki Itoh and Sanjay V. Malhotra (CRCPress)

Paper-XVI :DSE-4B(II)
Applied Organic Chemistry

Total Credits: 4
Contact hrs: 60

Unit I.

1. Theory of binary mixture analysis [08]

Types of organic compounds, nature and types of binary mixtures.
Reactions of acid, base, phenol and neutrals with sodium bicarbonate,
sodium hydroxide and hydrochloric acid
Principle of binary mixture separation.
Determination of type of the mixture
Separation of mixture- using aqueous medium and ether.

2. Green Chemistry [07]

Introduction
Twelve principles of green chemistry
Zeolites as green catalysts
Ultrasound assisted reactions
Reactions in ionic liquids
Solvent free reactions

Unit II:

3. Chemistry of cosmetics [15]

A general study including preparation and uses of- Hair dye, hairspray, shampoo, suntan lotions, face powder, lipsticks, talcum powder, nail enamel, cold creams, vanishing creams and shaving creams

Unit III:

4. Chemistry of perfumes [15]

A general study including preparation and uses of- antiperspirants, and artificial flavours
Essential oils and their importance in cosmetic industry with reference to Eugenol, geraniol, sandalwood oil, eucalyptus oil, rose oil, 2-phenyl ethyl alcohol, jasmone, civetone and muscone

Unit IV:

5. Fermentation [06]

Aerobic and anaerobic fermentation
Production of antibiotics-streptomycin
Production of vitamins-Vit.B12

6. Textile Chemistry [09]

Introduction, classification of fibers
Sizing: object of sizing, sizing ingredients and their functions
General idea of processes : singeing, desizing, scouring
Bleaching: Brief study of the outline of the process of bleaching cotton and synthetic material.
Dyeing : Study of dyeing of cellulosic material and synthetic fibers with dyes like direct, vat, reactive and dispersed dyes.

Reference Books

1. Industrial chemistry : B. K. Sharma(Goel Publishing House, Meerut)
2. Engineering Chemistry: P. C. Jain and M. Jain(Dhanpatrai and sons, Delhi)
3. Practical Organic Chemistry: A. I. Vogel
4. Advances in green chemistry-Chemical synthesis using Microwaveirradiation: R. S.Verma
5. A book of textile chemistry: A. J.Hall
6. Bleaching and Dyeing : Dr. V. Shenai
7. Sizing : D. B. Ajgaonkar
8. Chemical process industries : Shreve and Brinik (Ostin McGraw HillPublication, New York)
9. Medicinal and Pharmaceutical Chemistry: Hakishan, V. K. Kapoor(Vallabh Prakashan Pimpura New Delhi)
10. Industrial Chemistry, Vol. I:E. Stocchi (Ellis Horwood Ltd,UK)

PRACTICALS

- N.B. i. Use of Electronic balance with 0.001g accuracy is mandatory.
ii. Use of Scientific calculator is allowed.

Physical Chemistry

I) Non instrumental Experiments (Any Five) :

1. To determine the equilibrium constant of the reaction, $KI + I_2 \rightleftharpoons KI_3$ by the distribution method.
2. To determine the partition coefficient of CH_3COOH between H_2O and CCl_4 .
3. The study of energy of activation of first order reaction i.e. hydrolysis of methylacetate in presence of 0.5 N HCl.
4. The study of energy of activation of first order reaction i.e. hydrolysis of methylacetate in presence of 0.5 N H_2SO_4 .
5. The study of energy of activation of second order reaction i.e. reaction between $K_2S_2O_8$ and KI (Equal concentrations).
6. The study of energy of activation of second order reaction i.e. reaction between $K_2S_2O_8$ and KI (Unequal concentrations).
7. To study the hydrolysis of methyl acetate by using its two concentrations in presence of 0.5 N HCl and hence find velocity constant of the reaction.
8. To study the effect of addition of electrolyte (KCl) on the reaction between $K_2S_2O_8$ and KI (Equal concentrations).

II. Instrumental experiments

A. Potentiometry (Any Three).

1. Titration of strong acid with strong alkali.
2. Preparation of buffer solution and determination of their pH (Any five buffer solutions), -Theoretical calculation of pH values by using Henderson's equation.
3. Determination of standard electrode potential of Zn/Zn^{++} , Cu/Cu^{++} , Ag/Ag^+ (Any two).
4. Determination of solubility and solubility product of AgCl.
5. Titration of ferrous ammonium sulphate using $K_2Cr_2O_7$ solution and to calculate redox potential of Fe^{++} , Fe^{+++} system

B. Conductometry(any three).

1. Titration of weak acid with strong alkali.
2. Titration of a mixture of weak acid and strong acid with strong alkali.
3. To study the effect of substituent on dissociation constant of weak acid with respect to acetic acid and monochloro acetic acid (cell constant to be given).
4. To determine the velocity constant of hydrolysis of ethyl acetate by NaOH solution by conductometric method.

Reference Books :

1. Findlay's Practical Physical Chemistry(Longman)
2. Advanced Practical Physical Chemistry by J. B. Yadav, Goel publishing house.
3. Practical Physical Chemistry by B. D. Khosla, V. C. Garg (R. Chand and Co.)
4. Systematic experimental Physical Chemistry by Rajbhoj, Chandekar (Anjali Publication)
5. Practical Physical Chemistry : Nandkumari, Kothari and Lavande.
6. Practical Physical Chemistry by Gurtu (S. Chand).

Inorganic Chemistry Practicals

I. Gravimetric Estimations(G).

N. B. : Any two experiments from G1 to G3 and any two experiments from G4 to G7

G1. Gravimetric estimation of iron as ferric oxide from the given solution containing ferrous ammonium sulphate, copper sulphate and free sulphuric acid.

G2. Gravimetric estimation of zinc as zinc pyrophosphate from the given solution containing zinc sulphate, ferrous ammonium sulphate and free sulphuric acid.

G3. Gravimetric estimation of barium as barium sulphate from the given solution containing barium chloride, ferric chloride and free hydrochloric acid.

G4. Gravimetric estimation of manganese as manganese ammonium phosphate from the given solution containing manganese sulphate, copper sulphate and free sulphuric acid.

G5. Gravimetric estimation of barium as barium chromate from the given solution containing barium chloride, ferric chloride and free hydrochloric acid.

G6. Gravimetric estimation of Aluminium as Aluminiumoxinate i.e. tris (8-hydroxyquinolino) aluminate (III) from a given solution containing potash alum, copper sulphate and free sulphuric acid.

G7. Gravimetric estimation of nickel as bis (dimethylglyoximate) nickel (II) from the given solution containing nickel sulphate, ferrous ammonium sulphate and free sulphuric acid.

[For the gravimetric experiments, stock solution should be given in the range of 10 to 15 cm and asked to dilute to 100 cm (or the stock solution should be given in the range of 20 to 30 cm and asked to dilute to 250 cm). Use 50 cm of this diluted solution for estimation.]

II. Inorganic Preparations (P): (any five).

N. B.—1. Calculations of % yield is expected.

2. After preparation, physico-chemical characterization is expected with 5(Five) marks weightage in terms of:

- a) Name of central metal ion
- b) Oxidation number of metal ion
- c) Nature of ligand
- d) Nature of bonding
- e) Type of hybridization
- f) Inner orbital or outer orbital complex
- g) Geometry of the complex with structure
- h) Magnetic property of the compound
- i) Color of the compound
- j) Nature: Crystalline /Amorphous P1.

P1. Preparation of potassium trioxalatoferate(III)

P2. Preparation of potassium trioxalatoaluminate (III)

P3. Preparation of tris(ethylenediamine)nickel (II) thiosulphate

P4. Preparation of sodium hexanitrocobaltate (III)

P5. Preparation of ammonium diamminetetraethiocyanatochromate(III) (Reineck's salt)

P6. Preparation of nickel ferrite.

P7. Preparation of ammonium hexachlorostannate(IV)

P8. Preparation of tris(thiourea)cuprous(I) sulphate

III) Titrimetric Estimations:

A) Percentage Purity (any three)

V1. Determination of percentage purity of ferrous ammonium sulphate.

V2. Determination of percentage purity of tetramminecopper (II) sulphate.

V3. Determination of percentage purity of potassium trioxalatoaluminate(III).

V4. Determination of percentage purity of potassium trioxalatoferate (III).

Reference Books:

1. A text book of quantitative Inorganic Analysis - A. I. Vogel.
2. Text book of Quantitative Inorganic Analysis - Kolthoff and Sandell.
3. Experimental Inorganic Chemistry - Palmer W.G.
4. Advanced Practical Inorganic Chemistry - Adams and Raynor.
5. Handbook of Preparation Inorganic Chemistry. Vol. 1 and 11 -Brauer.
6. Manual in Dairy Chemistry - I.C.A.R. Sub-Committee on Dairy Education.
7. Chemical methods for environmental analysis - R. Ramesh and M. Anbu.

Organic Chemistry Practicals

I) Qualitative analysis

Separation of binary mixture and Identification of its components. 5g of mixture is to be given for separation. At least **08 mixtures** are to be separated.

Nature 1) Solid - Solid: 4 mixtures

2) Solid - Liquid : 2 mixtures

3) Liquid - Liquid : 2 mixtures

1) Solid - Solid Mixtures:

One mixture from each of the following types should be given:

i) Acid+Phenol ii) Acid +Base

iii) Acid+Neutral iv) Phenol +Base

v) Phenol+Neutral vi) Base +Neutral

2) Solid – Liquid Mixtures

One mixture of type Neutral + Neutral and One mixture of type Acid + Neutral should be given.

3) Liquid – Liquid Mixtures

One mixture of type Neutral + Neutral and One mixture of type Base + Neutral should be given.

Following compounds should be used for preparation of mixtures:

Acids: Benzoic acid, Phthalic acid, Salicylic acid, Cinnamic acid, Aspirin, Oxalic acid.

Phenols: α -naphthol, β -naphthol

Bases: o -nitroaniline, m-nitroaniline, p-nitroaniline, aniline, o-toluidine and N, N-dimethyl aniline.

Neutrals: Naphthalene, acetanilide, m-dinitrobenzene, chloroform, carbon tetrachloride, acetone, nitrobenzene, ethyl acetate, ethyl benzoate, acetophenone, bromobenzene, urea and thiourea.

II) Organic Preparations : (any four)

N.B.: a) Calculation of percentage practical yield.

b) Recrystallisation of crude product and its melting point.

c) The purity of the product may be confirmed by TLC.

1) Preparation of m-nitroaniline from m-dinitrobenzene.

2) Preparation of aspirin from salicylic acid.

3) Preparation of nerolin from β -naphthol.

4) Preparation of p-iodonitrobenzene from p-nitroaniline.

5) Preparation of benzene azo- β -naphthol.

6) Preparation of benzoic acid from cinnamic acid.

IV Preparation of Derivatives:

N.B.: During practical course, name of the organic compound should not be given.

1) Bromo derivative of aniline and cinnamic acid.

2) Nitro derivative of salicylic acid and nitrobenzene.

3) Benzoyl derivative of β -naphthol and aniline

4) Picrate derivative of anthracene and β -naphthol.

5) Oxalate and nitro derivatives of urea.

6) Anhydride derivative of phthalic acid.

7) Oxime derivatives of Ketones : Acetone and acetophenone.

8) 2: 4 DNP of acetophenone.

Reference Books:

1. Practical Organic Chemistry by A. I. Vogel.

2. Hand book of Organic qualitative analysis by H. T. Clarke.

3. A laboratory Hand Book of Organic qualitative analysis and separation by V. S. Kulkarni. Dastane Ramchandra & Co.

4. Practical Organic Chemistry by F.G. Mann and B.C. Saunders. Low-priced Text Book. ELBS. Longman.

5. Experiments in General Chemistry by C. N. R. Rao. Affiliated East-West Press Pvt. Ltd. Delhi.

6. Advanced Practical Organic Chemistry by N. K. Vishnoi. Vikas Publishing House Private Limited.
7. Comprehensive Practical Organic Chemistry Qualitative Analysis by V.K. Ahluwalia, Sunita Dhingra. University Press. Distributor - Orient Longman Ltd.
8. Comprehensive Practical Organic Chemistry Preparation and Quantitative Analysis by V.K. Ahluwalia, Renu Agarwal. University Press. Distributor-Orient Longmann Ltd.
9. Practical Chemistry-Physical-Inorganic-Organic and Viva-voce by Balwant Rai Satija. Allied Publishers Private Limited. 30
10. College Practical Chemistry by H. N. Patel, S. R. Jakali, H. P. Subhedar, Miss. S. P. Turakhia. Himalaya Publishing House, Mumbai.
11. College Practical Chemistry by Patel, Jakali, Mohandas, Israney Turakhia. Himalaya Publishing House, Mumbai.
12. Practice of thin layer chromatography by Joseph C. Touchstone, Murrell F. Dobbins. A Wiley - Interscience Publication John-Wiley & Sons.

Analytical Chemistry Practicals

Physical Chemistry Section

- Critical Solution Temperature: To determine the CST for phenol – water system.
- To determine the percentage composition of unknown mixture by (i) graphical method and (ii) by composition law (Densities of pure liquids A & B be given) refractometrically.
- To determine the molar refractivity of methyl acetate, ethyl acetate, n-hexane and carbon tetrachloride and calculate the refraction equivalents of C, H and Cl atoms refractometrically.
- To verify Lambert - Beer's law using CuSO_4 solution colorimetrically.
- To estimate Fe^{+++} ions by thiocyanate method colorimetrically.
- To estimate Fe^{+++} ions using salicylic acid by colorimetric titration colorimetrically.
- To determine the dissociation constant of monobasic acid (Acetic acid) pHmetrically.
- To determine the dissociation constant of dibasic acid (Malonic acid) pHmetrically.

Inorganic Chemistry Section:

- Determination of percentage of magnesium in the given sample of talcum powder by volumetric estimation.
- Determination of amount of aluminium in the given solution of potash alum by volumetric estimation.
- Determination of titrable acidity in the given sample of milk or lassi by volumetric estimation.
- Determination of Chemical Oxygen Demand of the given sample of industrial effluent by dichromate method by volumetric estimation.
- Determination of percentage purity of boric acid using supplied sodium hydroxide by volumetric estimation.
- Determination of amount of sodium present in the given solution of common salt using cation exchange resin (By Acid Base titration).
- Determination of amount of magnesium and zinc in the given solution containing (Mg^{++} and Zn^{++}) using anion exchange resin and standard solution of EDTA (By complexometric titration).

Organic Chemistry Section

- To estimate the amount of sucrose in given solution using Fehling's solution.
- To estimate amount of nitro group from the solution of m-nitroaniline
- To determine sap value of oil.
- To estimate the amount of acid and amide present in the given mixture of acid and amide
- To estimate the amount of acid and ester present in the given mixture of acid and ester
- To estimate the amount of formaldehyde from given formalin solution

Reference Books:

- College Practical Chemistry by H. N. Patel, S. R. Jakali, H. P. Subhedar, Miss. S. P. Turakhia. Himalaya Publishing House, Mumbai.
- College Practical Chemistry by Patel, Jakali, Mohandas, Israney Turakhia. Himalaya Publishing House, Mumbai.
- Experimental Inorganic Chemistry - Palmer W.G.
- A text book of quantitative Inorganic Analysis - A. I. Vogel.
- Text book of Quantitative Inorganic Analysis - Kolthoff and Sandell.
- Findlay's Practical Physical Chemistry (Longman)
- Practical Physical Chemistry by B. D. Khosla, V. C. Garg (R. Chand and Co.)
- Advanced Practical Physical Chemistry by J. B. Yadav, Goel publishing house.